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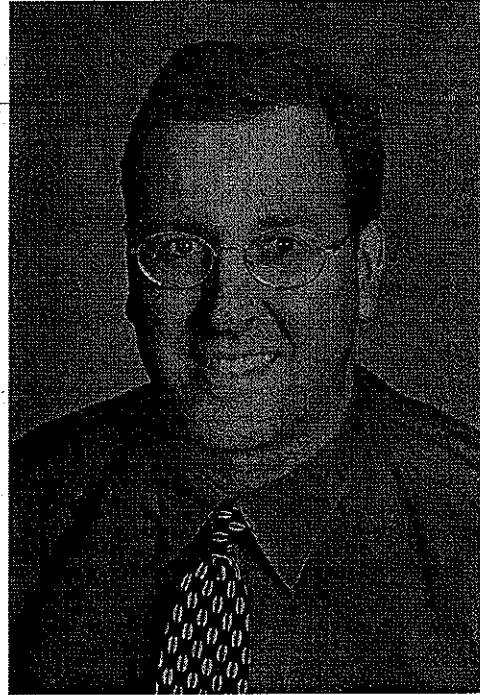
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Chris is a member in the firm's Boston office, practicing in the Business and Finance Section. He has extensive experience with commercial law, pharmaceutical and biotechnology licensing and joint ventures, food and drug regulatory matters and dispute resolution.

Prior to joining the firm, Chris held various positions at AstraZeneca, a global pharmaceutical company, where his responsibilities included negotiating and drafting license agreements for drug compounds and related technology and a broad range of agreements facilitating drug discovery, product development, and sales and marketing efforts.

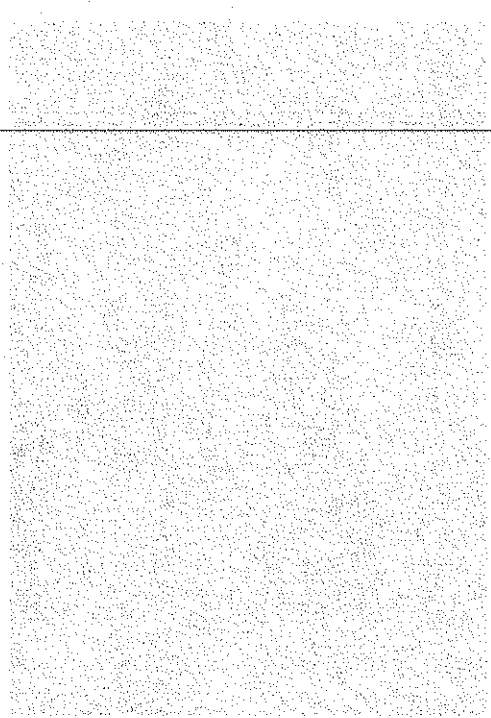
In addition, while employed by another Boston-area law firm, Chris handled general corporate and securities matters. He provided general legal advice and services to businesses including shareholder relations, credit and financial matters, strategic planning, and corporate governance issues.

Chris is admitted to practice in Massachusetts. He received his B.A. *cum laude* from Connecticut College in 1986, where he was a member of Phi Beta Kappa. He was awarded his J.D. *cum laude* from Boston College Law School in 1991.



Christopher J. Denn

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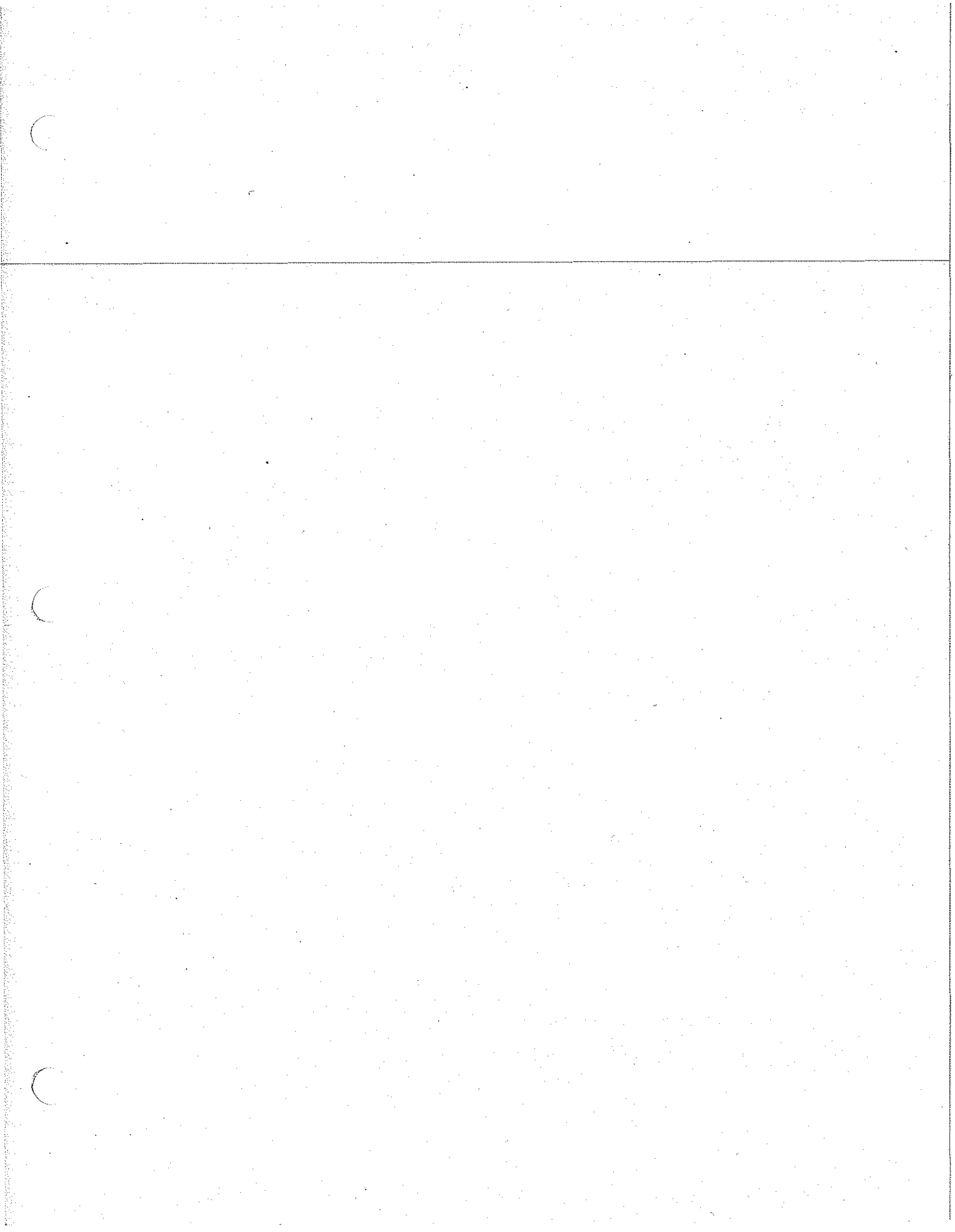


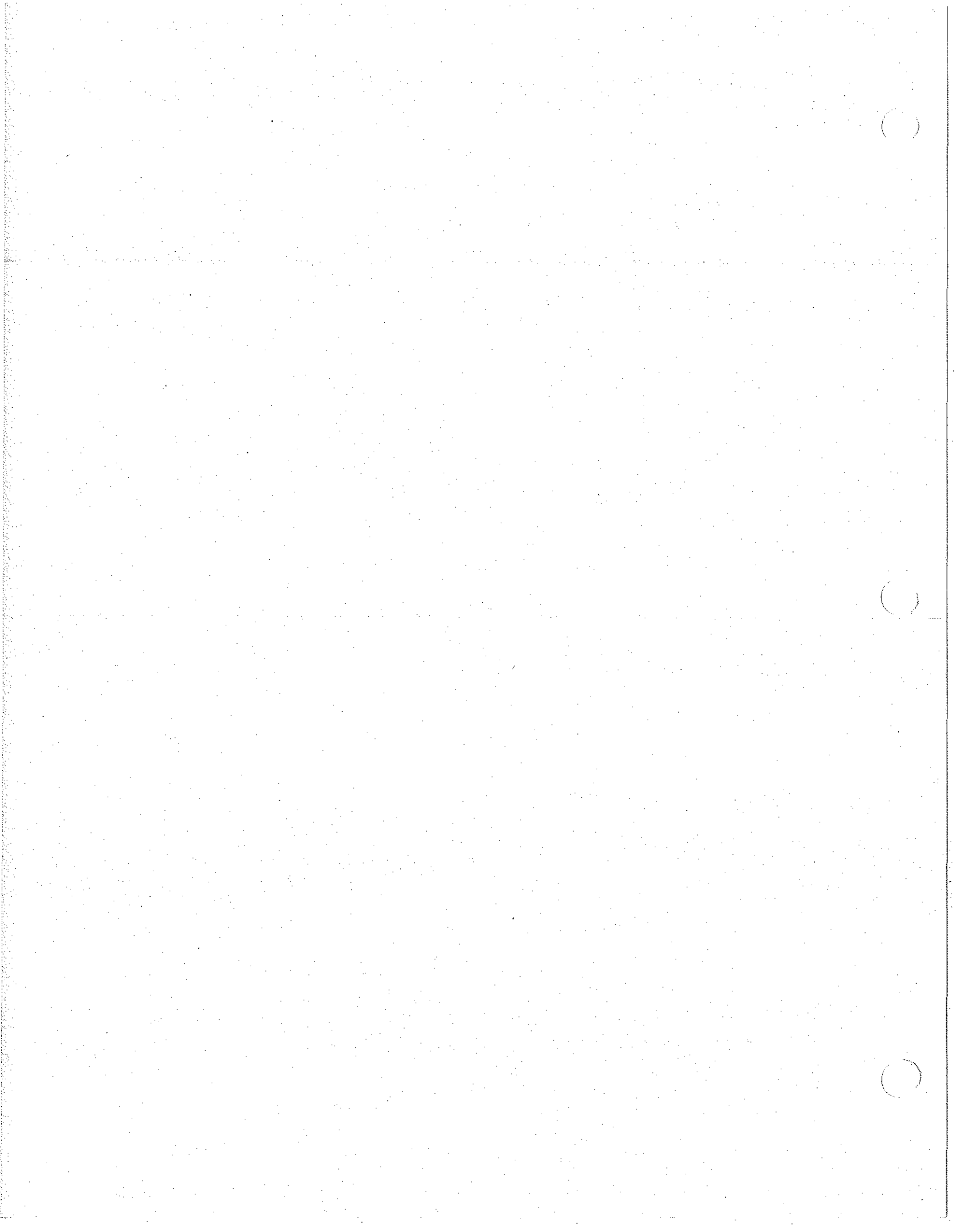
The first part of the report is a summary of the work done during the year. It covers the general progress of the project and the results of the various experiments. The second part is a detailed description of the methods used in the experiments. This includes a description of the apparatus used and the procedures followed. The third part is a discussion of the results of the experiments. This includes a comparison of the results with those obtained in previous work and a discussion of the implications of the results.

The results of the experiments show that the rate of reaction is proportional to the concentration of the reactants. This is in agreement with the theoretical prediction. The activation energy of the reaction has been determined to be 15 kcal/mole. This value is in good agreement with that reported in the literature. The order of reaction with respect to the reactants is 1.5 and 1.0 respectively. This is also in agreement with the theoretical prediction.

The rate of reaction is also affected by the presence of a catalyst. The rate of reaction is increased by a factor of 10 when a catalyst is present. This is in agreement with the theoretical prediction. The activation energy of the reaction is lowered by 5 kcal/mole when a catalyst is present. This is also in agreement with the theoretical prediction.

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- Milestones
- Compensation for services
- Earned royalty

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- **Royalty rate**
- **Royalty reductions**
- **Royalty period**
- **Minimum royalty**
- **Payments from sublicensees**

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- Patent Expiration
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Other Issues

Hart –Scott– Radino

- Size of persons
- Size of transaction
- Notification
- Waiting period and early termination

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